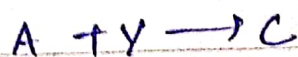
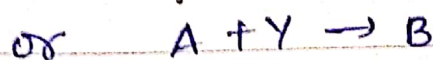


## Types of Composite Reactions

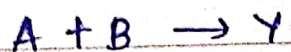
### ① Simultaneous Reactions



Two products can be obtained from same set of reactants

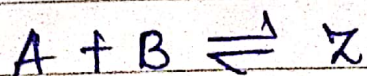


Sometimes there may be reactant competition.



where B and C compete with A for reaction

### ② Opposing reactions



when the reaction takes place in both forward and backward directions

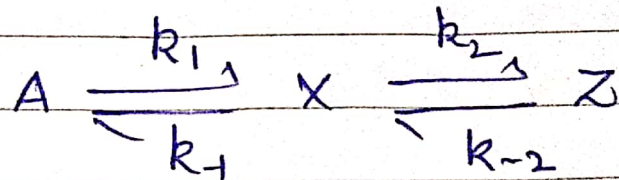
### ③ Consecutive reactions:



When the reaction proceeds with multisteps before final product formation

## Rate equation for composite reaction:

For composite reaction



There can be four elementary reactions and their rate is given as

$$V_1 = k_1 [A]$$

$$V_{-1} = k_{-1} [X]$$

$$V_2 = k_2 [X]$$

$$V_{-2} = k_{-2} [Z]$$

for intermediate X total rate of ~~formation~~ <sup>formation</sup> ~~consumption~~ <sup>consumption</sup>

$$V_x = V_1 + V_{-2}$$

$$= k_1 [A] + k_{-2} [Z]$$

and total rate of X consumption is

$$V_{-x} = V_{-1} + V_2$$

$$= k_{-1} [X] + k_2 [X]$$

$$= [(k_{-1}) + k_2] [X]$$

$$= (k_{-1} + k_2) [X]$$

Now for consecutive reaction we have



for if initial concentration of [A] is  $[A]_0$ . The rate of reaction for A is

$$V_1 = \frac{d[A]}{dt} = -k_1[A]$$

Integration of this term in boundary condition  $[A] = [A]_0$  for  $t = 0$

$$[A] = [A]_0 e^{-k_1 t}$$

For the rate of formation of [X]

$$\frac{d[X]}{dt} = k_1[A] - k_2[X]$$

$$= k_1[A]_0 e^{-k_1 t} - k_2[X]$$

On integrating we get

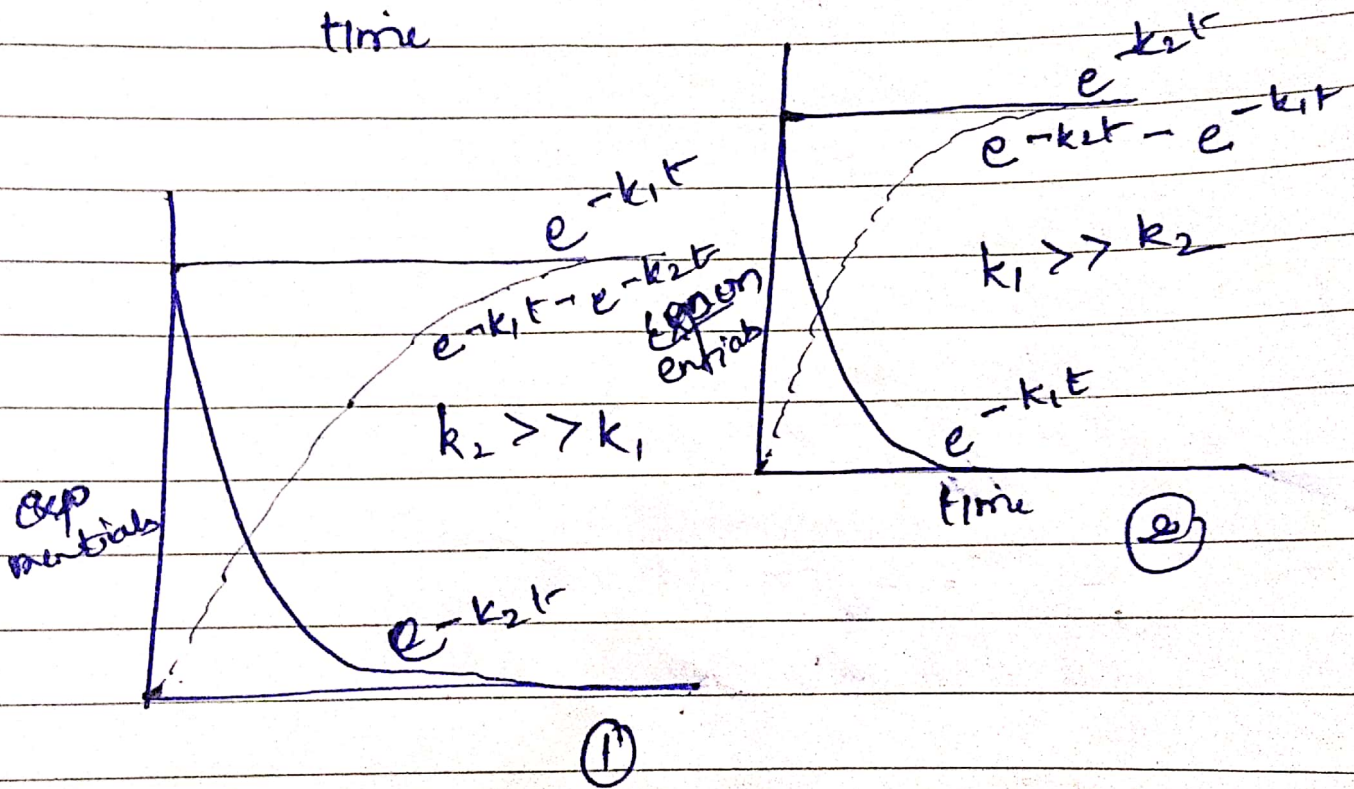
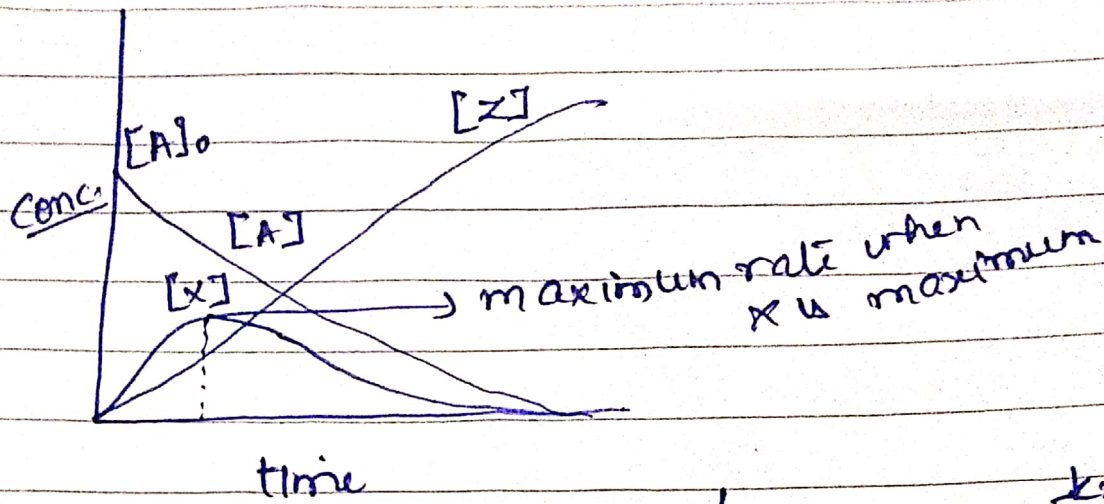
$$[X] = [A]_0 \frac{k_1}{k_2 - k_1} (e^{-k_1 t} - e^{-k_2 t})$$

$$[A] + [X] + [Z] = [A]_0$$

so  $[Z] = [A]_0 - [A] - [X]$

$$[Z] =$$

$$\frac{[A]_0}{k_2 - k_1} [k_2(1 - e^{-k_1 t}) - k_1(1 - e^{-k_2 t})]$$



The graph shows that formation of  $[Z]$  is proportional to the concentration of  $X$ . The rate is zero when  $[X] = 0$  and reaches maximum for maximum value of  $[X]$ .

Graph (1) and (2) show two limiting conditions.